

Errata

E. Cesarotti, R. Ugo and L. Kaplan, A discussion of the different kinds of solute–solute and solute–solvent interactions acting in homogeneous catalysis by transition metal complexes. *Coord. Chem. Rev.*, 43 (1982) 275–298.

p. 290, below Table 7, should read:

In general, the direction of the effect on the rate of a reaction of certain changes in the nature or magnitude of solute–solute and solute–solvent interactions is dependent upon the structural characteristics and charge distribution in the reactants relative to the transition state(s) of the rate-determining step(s); such factors can vary significantly as a result of seemingly minor changes in the composition of a chemical system. It is therefore impossible to make generally applicable statements regarding the direction of the effect of certain changes in, e.g. the nature of the solvent or of the counter-ion, on the rate of this particular type of reaction. Diverse trends are therefore to be expected *.

* Note that, although the results above involving crown ethers, and other results in the literature fall into the general pattern in a way consistent with the idea that a crown-complexed cation forms a weaker ion pair than the 'uncomplexed' (i.e. solvent-complexed) cation, this need not have been the outcome a priori[†] (p. 291). Therefore, it would be inappropriate to conclude, for example, solely from the fact that addition of crown ether decreases the rate of a reaction, that the reaction is 'assisted by ion pairing'.

p. 291, line 22 should end with a full stop; line 23 should be deleted. The footnote should be indicated by a dagger and lines 10 and 11 of the footnote should read:

specific, close contact with K^+ , as methanol's interaction is), does not suffer when K^+ becomes K^+ (crown) as much as does that between K^+ and methanol. (b) Below -10°C in THF the ion-pair

p. 292, $\text{CH}_3\text{CO}_2\text{CH}_3$ on the 3rd line of the heading of Table 8 follows on after the + at the end of the second line.

The first complete paragraph after Table 8 should read:

It is clear that ion-pairing effects are very relevant, since the complex mechanism of reaction may involve steps such as carbon monoxide insertion, ligand exchange, etc. which are known to be sensitive to ion pairing [see sections (i), (ii) and (iii)].